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Following the indium triorganyls and their adducts with Lewis bases in Section 1, the broad field of compounds of the general type R In- n 3 n (n = 1, 2) is treated in sections 2 to 9; X represents a ligand bonded with a non-carbon atom to the indium atom.

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compounds n 1 to 4 close the series of trivalent n 4 organoindium compounds and are collected in section 11 compounds of formally low valent indium (In<sup>+</sup>, In<sub>2</sub> and In<sub>3</sub>), with one R 1InR species having an In-In bond, form Section 12; 2 2 an extended chapter therein is dedicated to the young area of Cp<sup>+</sup>In compounds in which i formalln is coordinated in an T] 5 manner.

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Part of the Gmelin Handbook of Inorganic and Organometallic Chemistry / Gmelin Handbuch der Anorganischen Chemie book series (GMELIN, volume I-n / 1- / 1) Abstract The compounds in this section are mostly thiolates, R 2 InSR<sub>2</sub> and are described in Table 46.

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Boron has all the best tunes. That may well be the first impression of the Group 13 elements. The chemical literature fosters the impression not only in the primary journals, but also in asteady outflowbooks focussing more or less closely on boron and its compounds. The same preoccupation with boron is apparent in the coverage received by the Group 13 elements in the comprehensive and regularly updated volume of the Gmelin Handbook. Yet such an imbalance cannot be explained by any inherent lack ofvariety, interest or consequence in the 'heavier elements. Aluminium is the most abundant metal in the earth's crust; in the industrialised world the metal is second only to iron in its usage, and its compounds can justifiably be said to touch our lives daily - to the potential detriment of those and other lives, some would argue. From being chemical curios, gallium and indium have now gained considerably prominence as sources of compound semiconductors like gallium arsenide and indium antimonide. Nor is there any want ofincident in the chemistriesofthe heavier Group 13 elements. In their redox, coordination and structural properties, there is to be found music indeed, notable not always for its harmony but invariably for its richness and variety. Thisbook seeks to redress the balance with a definitive, wide-rangingand up-to-date review of the chemistry of the Group 13 metals aluminium, gallium, indium and thallium.

The present volume in the organogermanium series describes mononuclear compounds containing only germanium-carbon and germanium-hydrogen bonds (Chapter 1.3). Germanium hydrides with other additional non-carbon ligands, such as halogen or oxygen bonded groups, appear in later chapters according to the Gmelin principle of the last posi tion. Compounds with Ge-H and Ge-O bonds have already been described in Volume 5, Section 1.5.1.4, pp. 50/62. The present volume covers the literature to the end of 1992 and includes many references up to 1994. The nomenclature recommended by IUPAC has been generally adhered to. However, compound names were largely avoided, as most of the compounds are presented in tables and are only identified by their formulas. Many of the data in the tables appear in abbreviated form without units; general explanations are given on pp. X /X I. The volume contains an empirical formula index (p. 327) and a ligand formula index (p.341). The editor wishes to express his gratitude to the former author, Professor J. E. Drake, and to Professor J. Satge for his kind advice and fruitful collaboration. Thanks are due also to Dr. A. R. Pebler for editing the English text and to Mr. H.-G. Karrenberg for drawing the numerous formulas and molecular structures.

This volume starts the updating of the information on organonickel chemistry, one of the most active areas of modern organometallic chemistry. It covers all mononuclear compounds containing one nickel-to-carbon bond. The main part is devoted to the family of planar NiII compounds, RNi(D)2X, which within 20 years has increased from about 250 to more than 1000 members. This research interest has been stimulated by the fact that many compounds insert various unsaturated species into the Ni-C bond to create new organonickel species. Moreover, these compounds are important as intermediates in many catalytic processes involving organic halides. Compounds containing h<sup>2</sup>-bonded heteroalkene molecules including the C1 fragments CO2 and CS2, and related metallocyclic compounds are also emphasized in this volume. These types of compounds play an important role in a number of nickel-catalyzed co-oligomerization reactions involving unsaturated organic compounds. Most of the compounds are well-characterized by modern spectroscopic techniques. The various types of compounds are in most cases confirmed by X-ray diffraction. The author of these volumes, a member of the M them group, has made many contributions of his own and is eminently qualified to critically review this area of research.

This volume 8 is the fourth in a series dealing with organomolybdenum compounds. An Empirical Formula Index and a Ligand Formula Index provide ready access to the compounds covered. Volume 5 describes mononuclear organomolybdenum compounds with isocyanide, 3 4 carbene, carbyne, alkynyl, alkene, alkyne, L, and L ligands with and without additional CO groups. Volume 6 starts the description of mononuclear organomolybdenum compounds with 5 one L Ligand, a Ligand bonded to molybdenum by five carbon atoms. The compounds contain either zero or one CO group bonded to the molybdenum atom. Volume 7 continues the 5 description of L- molybdenum compounds containing two CO groups, but no additional n] 5 ligands. This volume describes L-molybdenum compounds with two CO groups and additional 1 4 L to L ligands. Following the nomenclature used in this series of organomolybdenum com pounds, n] is an organic Ligand bonded by n C atoms to molybdenum, and mo is an electron 2 donor Ligand with m donor electrons. Thus 0 denotes a Ligand such as PR<sub>3</sub>. Many of the data, particular those in tables, are given in an abbreviated form without units; for explanations see p. X. Additional information, if necessary, is given before the individual table. Frankfurt am Main Manfred Winter November 1992 Wolfgang Petz X Remarks on Abbreviations and Dimensions Many compounds in this volume are presented in tables in which numerous abbreviations are used, the dimensions are omitted for the sake of conciseness. This necessitates the following clarifications.

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